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### Edexcel Gcse Chemistry Textbook Answers

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EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION B Chapter 7 1 a) b)  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$  c) Relights a glowing splint 2 a) non-metal (sulfur) b) metal (sodium) c) metal (copper) d) non-metal (carbon as graphite) 3 a) React with a named heated metal (most obvious would be to pass over heated

#### EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL ...

Chemistry EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION A Chapter 1 1 a) melting b) freezing c) subliming / sublimation d) subliming / sublimation 2 a) Note: Solids should have regularly packed particles touching Liquids ...

#### Z04 EDIC SB IGCSE 85169 ANS 345-374

346 ANSWERS g 82, lead h 6 Palladium is a metal and so is likely to have any of the following properties: • good conductor of electricity • forms a basic oxide • is shiny when polished or freshly cut • b is malleable • is ductile • is a good conductor of heat The first ...

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#### EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL ...

EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION D Chapter 18 1 a) (i) methane (ii) propane (iii) pentane (iv) propene (v) ethene (vi) but-1-ene b) 2 a) The existence of molecules with the same molecular formula but ...

#### Answers Section 1. Principles of chemistry

Answers Edexcel International GCSE Chemistry Teacher Pack 6 ©HarperCollinsPublishers Ltd 2012 Page 56 1 Fe 2 O 3 (010 moles of Fe reacts with 015 moles of O) 2 ZnO (020 moles of Zn combine with 020 moles of O) 3 C 4 H 10 (empirical formula mass is 29 - ...

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Edexcel Igcse Chemistry 9-1 Textbook Answerspdf - Free impression often given at GCSE, such compounds are very common - although in the great majority of cases, there are more than 8 electrons around one atom rather than fewer

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EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION C Chapter 16 1 Burn sulfur in air to give sulfur dioxide  $S(s) + O_2(g) \rightarrow SO_2(g)$  Pass this with more air over a vanadium(V) oxide catalyst at a temperature of 450°C and a pressure of 1

to 2 atmospheres to give sulfur trioxide

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